

Paper Reference(s) 1CH0/1F  
Pearson Edexcel Level 1/Level 2 GCSE (9–1)

Chemistry  
Paper 1  
Foundation Tier

Thursday 14 May 2020 – Morning

Time: 1 hour 45 minutes plus your additional time allowance

Diagram Booklet

In the boxes below, write your name, centre number and candidate number.

Surname					
Other names					
Centre Number					
Candidate Number					

**INSTRUCTIONS**

There may be spare copies of some diagrams in case you need them.

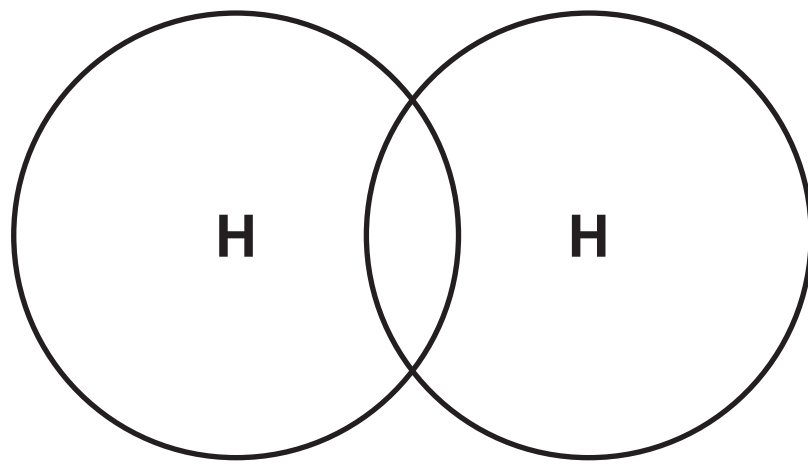
**THIS DIAGRAM BOOKLET *MUST* BE RETURNED WITH THE  
QUESTION PAPER AT THE END OF THE EXAMINATION.**

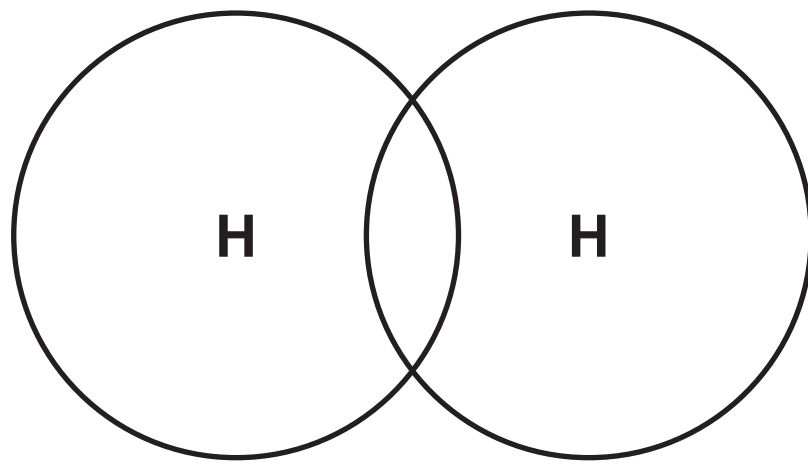
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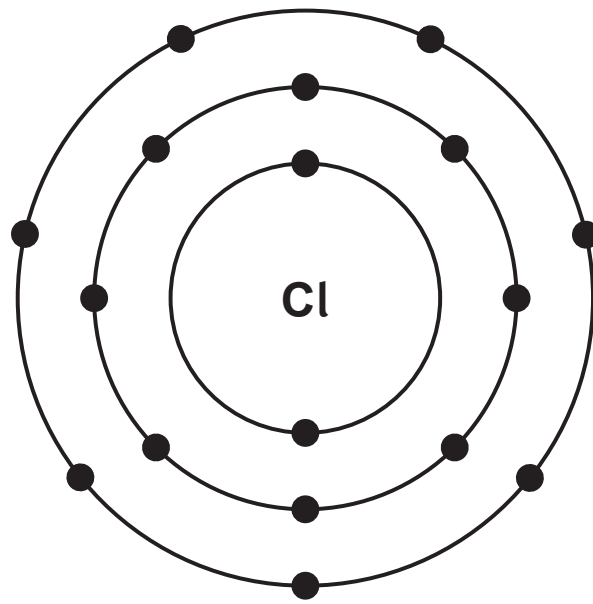
**Question 1(b)(ii)****Figure 1**

**Question 1(b)(iii)****Figure 2**

**Question 1(b)(iii) (Spare copy)****Figure 2**

## Question 2(a)

Figure 3



Question 2(b)

Figure 4

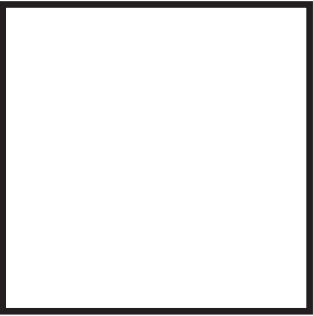
	relative mass	relative charge
proton	1	+1
neutron	1	0
electron	very small	-1



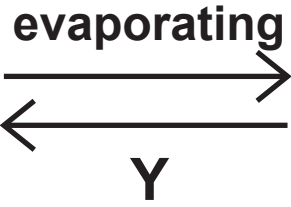
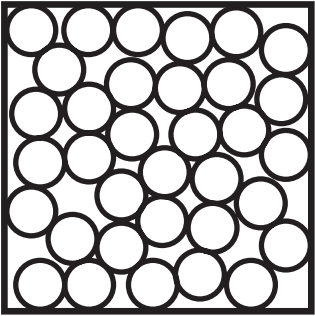
Question 4(d)

Figure 5

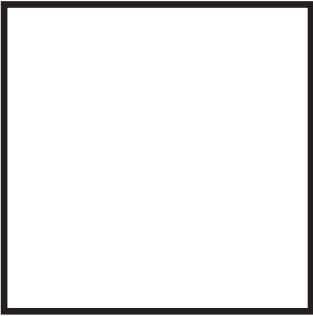
solid gallium



liquid gallium



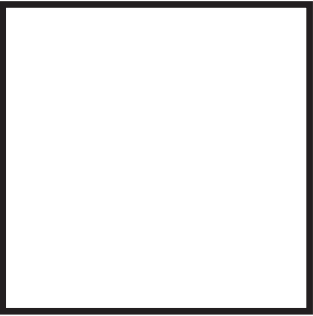
gaseous gallium



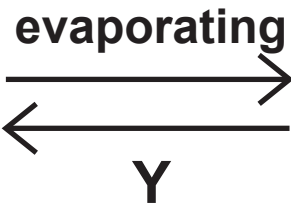
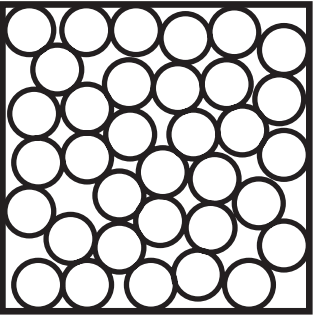
Question 4(d)

Figure 5

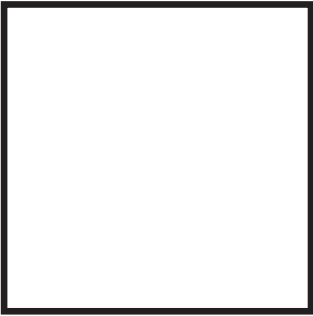
solid gallium



liquid gallium

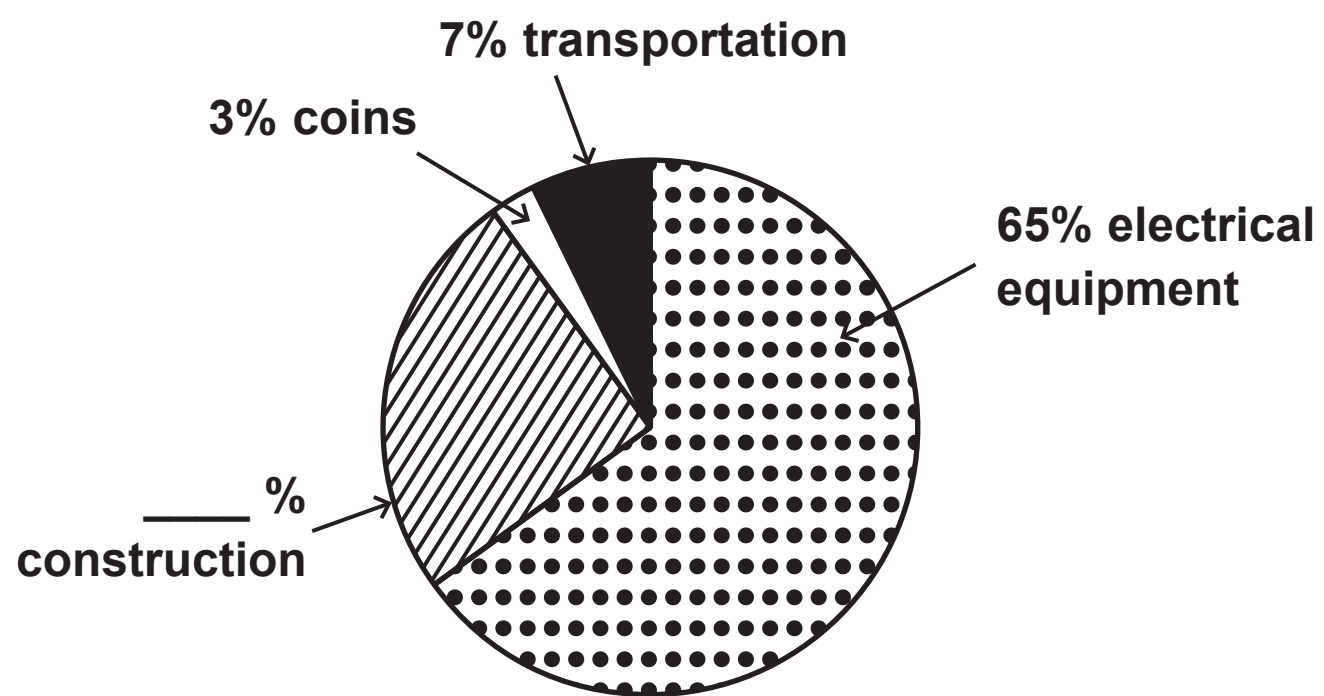


gaseous gallium



## Question 5(a)

Figure 6



Question 5(b)

Figure 7

	true	false
iron is a poor conductor of heat		✓
iron can act as a catalyst		
iron forms compounds that are coloured		
iron has a low density		
iron has a very high melting point		

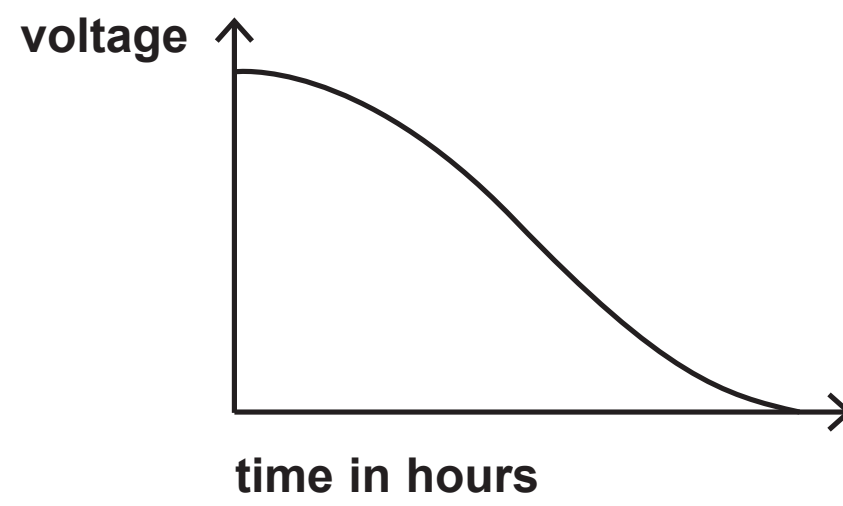
Question 5(b) (Spare copy)

Figure 7

	true	false
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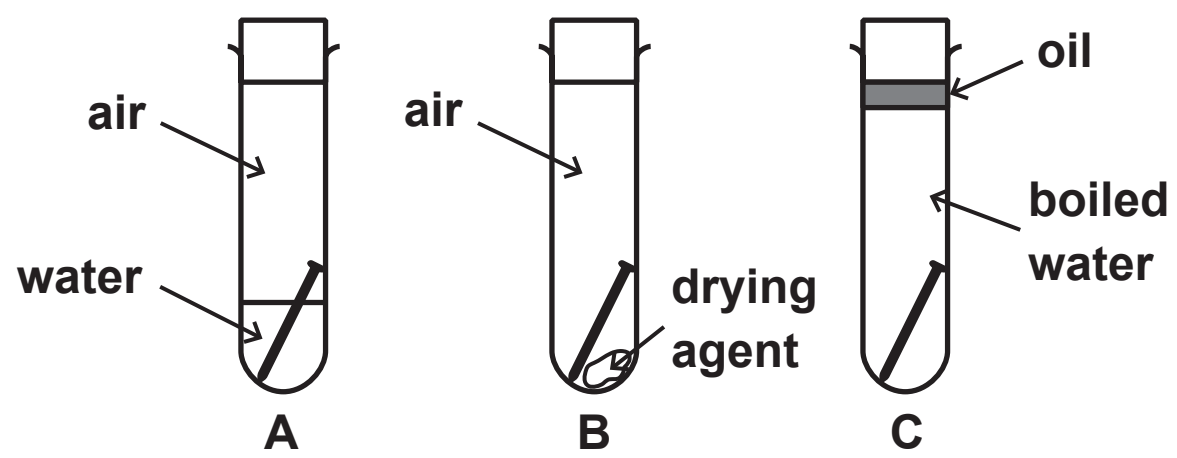
## Question 5(d)

Figure 8



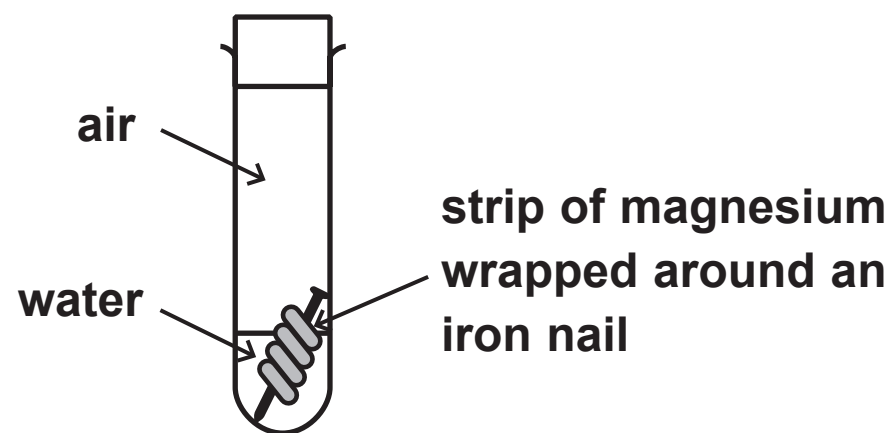
## Question 7(a)(i)

Figure 9



## Question 7(a)(ii)

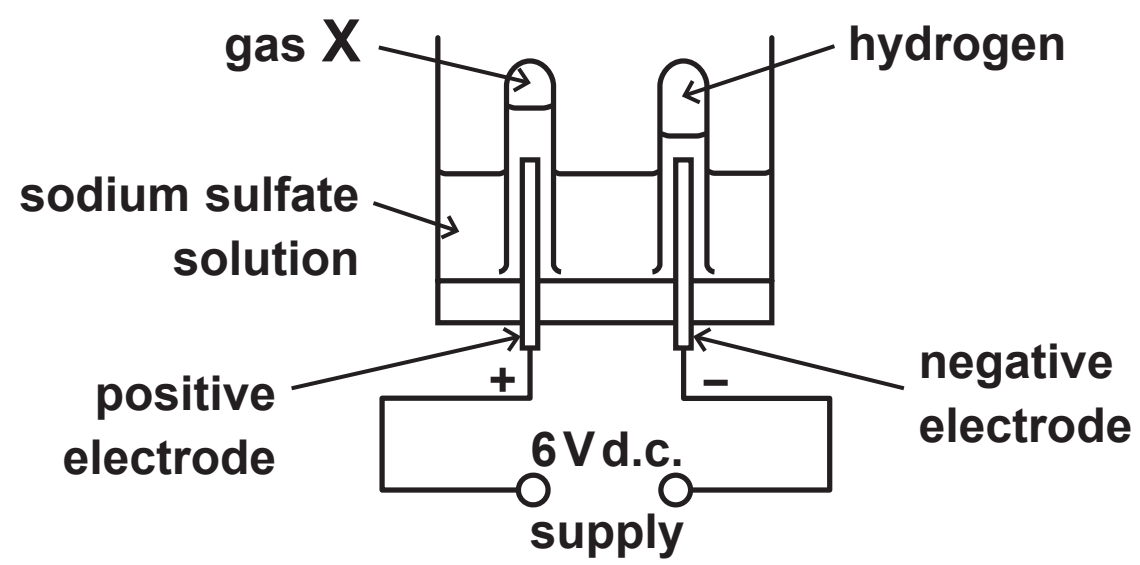
Figure 10





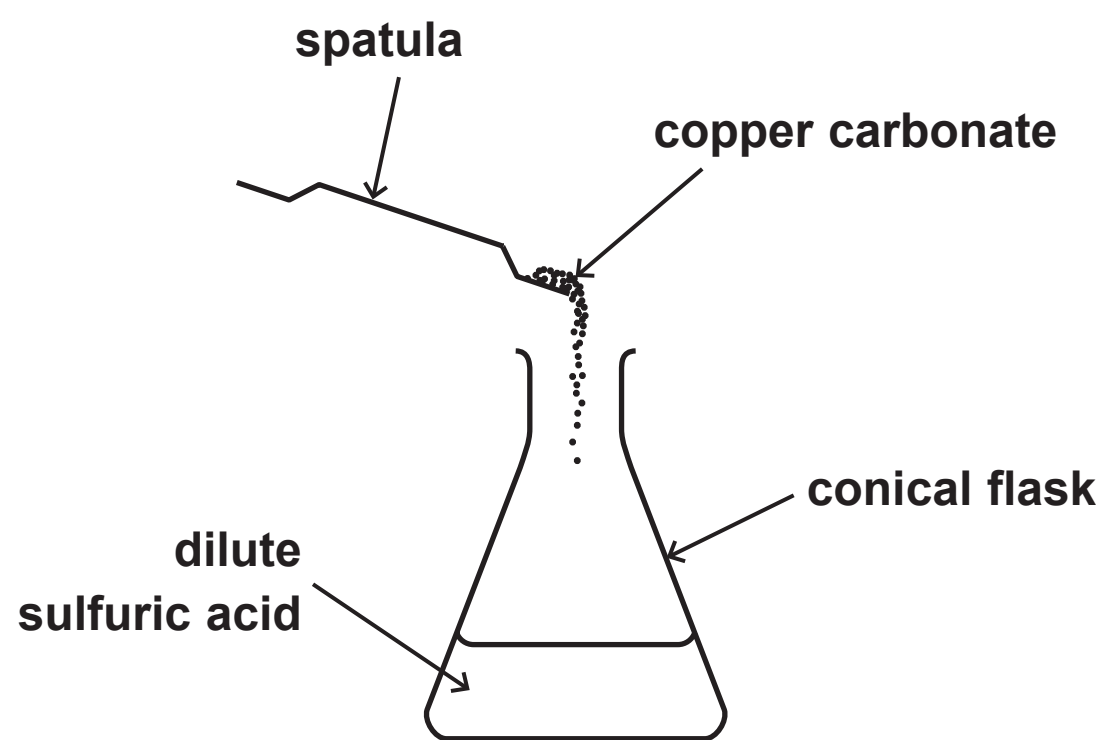
## Question 8

Figure 11



## Question 9(b)

Figure 12



**Question 10**

- STEP J** pour the potassium hydroxide solution into a conical flask and add a few drops of indicator to this solution
- STEP K** fill a burette with the dilute hydrochloric acid and record the initial reading from the burette
- STEP L** use a measuring cylinder to obtain  $25\text{cm}^3$  of potassium hydroxide solution
- STEP M** take a final reading from the burette and calculate the volume of the dilute hydrochloric acid reacted
- STEP N** run the dilute hydrochloric acid from the burette into the conical flask until the indicator changes colour